

## Periodic Trends Homework

- 1) Why does electronegativity increase as you move across the periodic table?
- 2) Why do elements in the same families tend to have similar electronegativities?
- 3) What's the difference between electronegativity and ionization energy?
- 4) Why does atomic radius decrease as you move across the periodic table?
- 5) What does the octet rule have to do with the periodic trends?
- 6) The first ionization energy of beryllium is 9.322 eV, the second ionization energy is 18.211 eV, and the third ionization energy is 153.893 eV. Explain why the third ionization energy of beryllium is so much higher than the first two.