

Properties of Ionic Compounds Homework

- 1) What charge will the following elements have when they follow the octet rule to become like the nearest noble gas?
 - a) bromine _____
 - b) nitrogen _____
 - c) magnesium _____
 - d) sulfur _____

- 2) Explain the process by which a neutral atom becomes a cation. Use aluminum as an example.

- 3) Why is LiF considered to be an ionic compound while GaAs is not? What does this have to do with their positions on the periodic table?

- 4) Explain why ionic compounds have high melting and boiling points.

- 5) What test that you could perform on a compound would be most important in determining if it were ionic or not? Explain your reasoning.

- 6) What is lattice energy?