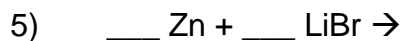
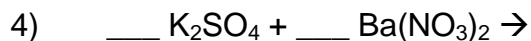
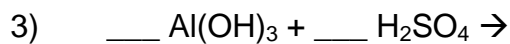


Review for Final #1

Translate the following statements into complete, balanced chemical equations:

- 1) When dissolved potassium iodide combines with a solution of lead (II) nitrate, lead iodide precipitate and dissolved potassium nitrate are formed.
- 2) Nitric acid undergoes a neutralization reaction with solid calcium hydroxide in an exothermic process.

Predict the products of the following reactions. Keep in mind that some of these reactions may not occur – you should indicate which will not and why.



- 6) For the reaction $\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$, indicate how many grams of hydrogen gas will be required to make 350 grams of ammonia. Assume that there is plenty of excess hydrogen.

- 7) The reaction from #6 is an equilibrium. What does this tell you about the validity of the calculation you performed above?